

red happens @
↓

oxid happens @
↓

$$E^{\circ}_{\text{cell}} = E^{\circ}_{\text{red}}(\text{cathode}) - E^{\circ}_{\text{red}}(\text{anode})$$

$$E^{\circ}_{\text{cathode}} [\text{Cu}^{2+} \rightarrow \text{Cu}(s)] = 0.342 \text{ V}$$

$$E^{\circ}_{\text{anode}} [\text{Zn}^{2+} \rightarrow \text{Zn}(s)] = -0.762 \text{ V}$$